



# University of Pretoria Yearbook 2023

## General chemistry 117 (CMY 117)

**Qualification** Undergraduate

**Faculty** Faculty of Natural and Agricultural Sciences

**Module credits** 16.00

**NQF Level** 05

**Programmes** BEd (Senior Phase and Further Education and Training Teaching)

BSc (Computer Science)

Bachelor of Dietetics [BDietetics]

BSc (Applied Mathematics)

BSc (Biochemistry)

BSc (Biological Sciences)

BSc (Biotechnology)

BSc (Chemistry)

BSc (Ecology)

BSc (Engineering and Environmental Geology)

BSc (Entomology)

BSc (Food Management) *Culinary Science*

BSc (Food Management) *Nutrition*

BSc (Food Science)

BSc (Genetics)

BSc (Geography) *Geography and Environmental Science*

BSc (Geology)

BSc (Human Genetics)

BSc (Human Physiology)

BSc (Human Physiology, Genetics and Psychology)

BSc (Mathematics)

BSc (Medical Sciences)

BSc (Meteorology)

BSc (Microbiology)



BSc (Physics)

BSc (Plant Science)

BSc (Zoology)

BScAgric (Agricultural Economics and Agribusiness Management)

BScAgric (Animal Science)

BScAgric (Applied Plant and Soil Sciences)

BScAgric (Plant Pathology)

**Service modules**

Faculty of Engineering, Built Environment and Information Technology

Faculty of Education

Faculty of Health Sciences

Faculty of Veterinary Science

**Prerequisites**

A candidate must have Mathematics for at least 60% and 60% for Physical Sciences.

**Contact time**

1 practical per week, 4 lectures per week

**Language of tuition**

Module is presented in English

**Department**

Chemistry

**Period of presentation** Semester 1

**Module content**

General introduction to inorganic, analytical and physical chemistry. Atomic structure and periodicity. Molecular structure and chemical bonding using the VSEOR model. Nomenclature of inorganic ions and compounds. Classification of reactions: precipitation, acid-base, redox reactions and gas-forming reactions. Mole concept and stoichiometric calculations concerning chemical formulas and chemical reactions. Principles of reactivity: energy and chemical reactions. Physical behaviour gases, liquids, solids and solutions and the role of intermolecular forces. Rate of reactions: Introduction to chemical kinetics.

**Regulations and rules**

The regulations and rules for the degrees published here are subject to change and may be amended after the publication of this information.

The [General Academic Regulations \(G Regulations\)](#) and [General Student Rules](#) apply to all faculties and registered students of the University, as well as all prospective students who have accepted an offer of a place at the University of Pretoria. On registering for a programme, the student bears the responsibility of ensuring that they familiarise themselves with the General Academic Regulations applicable to their registration, as well as the relevant faculty-specific and programme-specific regulations and information as stipulated in the relevant yearbook. Ignorance concerning these regulations will not be accepted as an excuse for any transgression, or basis for an exception to any of the aforementioned regulations.



### **University of Pretoria Programme Qualification Mix (PQM) verification project**

The higher education sector has undergone an extensive alignment to the Higher Education Qualification Sub-Framework (HEQF) across all institutions in South Africa. In order to comply with the HEQSF, all institutions are legally required to participate in a national initiative led by regulatory bodies such as the Department of Higher Education and Training (DHET), the Council on Higher Education (CHE), and the South African Qualifications Authority (SAQA). The University of Pretoria is presently engaged in an ongoing effort to align its qualifications and programmes with the HEQSF criteria. Current and prospective students should take note that changes to UP qualification and programme names, may occur as a result of the HEQSF initiative. Students are advised to contact their faculties if they have any questions.