



# University of Pretoria Yearbook 2024

## General chemistry 117 (CMY 117)

**Qualification** Undergraduate

**Faculty** Faculty of Natural and Agricultural Sciences

**Module credits** 16.00

**NQF Level** 05

**Programmes** BEd (Senior Phase and Further Education and Training Teaching)

*BSc Computer Science*

*Bachelor of Dietetics [BDietetics]*

*BSc (Biological Sciences)*

*BSc (Food Management) Culinary Science*

*BSc (Food Management) Nutrition*

*BSc (Geography) Geography and Environmental Science*

*BSc Applied Mathematics*

*BSc Biochemistry*

*BSc Biotechnology*

*BSc Chemistry*

*BSc Ecology*

*BSc Entomology*

*BSc Environmental and Engineering Geology*

*BSc Food Science*

*BSc Genetics*

*BSc Geology*

*BSc Human Genetics*

*BSc Human Physiology*

*BSc Human Physiology, Genetics and Psychology*

*BSc Mathematics*

*BSc Medical Sciences*

*BSc Meteorology*

*BSc Microbiology*



*BSc Physics*

*BSc Plant Science*

*BSc Zoology*

*BScAgric Agricultural Economics and Agribusiness Management*

*BScAgric Animal Science*

*BScAgric Applied Plant and Soil Sciences*

*BScAgric Plant Pathology*

**Service modules**

Faculty of Engineering, Built Environment and Information Technology

Faculty of Education

Faculty of Health Sciences

Faculty of Veterinary Science

**Prerequisites**

A candidate must have Mathematics for at least 60% and 60% for Physical Sciences.

**Contact time**

1 practical per week, 4 lectures per week

**Language of tuition**

Module is presented in English

**Department**

Chemistry

**Period of presentation** Semester 1

**Module content**

General introduction to inorganic, analytical and physical chemistry. Atomic structure and periodicity. Molecular structure and chemical bonding using the VSEPR-model. Nomenclature of inorganic ions and compounds. Classification of reactions: precipitation, acid-base, redox reactions and gas-forming reactions. Mole concept and stoichiometric calculations concerning chemical formulas and chemical reactions. Principles of reactivity: energy and chemical reactions. Physical behaviour gases, liquids, solids and solutions and the role of intermolecular forces. Rate of reactions: Introduction to chemical kinetics.

**General Academic Regulations and Student Rules**

The [General Academic Regulations \(G Regulations\)](#) and [General Student Rules](#) apply to all faculties and registered students of the University, as well as all prospective students who have accepted an offer of a place at the University of Pretoria. On registering for a programme, the student bears the responsibility of ensuring that they familiarise themselves with the General Academic Regulations applicable to their registration, as well as the relevant faculty-specific and programme-specific regulations and information as stipulated in the relevant yearbook. Ignorance concerning these regulations will not be accepted as an excuse for any transgression, or basis for an exception to any of the aforementioned regulations. The G Regulations are updated annually and may be amended after the publication of this information.

**Regulations, degree requirements and information**

The faculty regulations, information on and requirements for the degrees published here are subject to change



and may be amended after the publication of this information.

**University of Pretoria Programme Qualification Mix (PQM) verification project**

The higher education sector has undergone an extensive alignment to the Higher Education Qualification Sub-Framework (HEQSF) across all institutions in South Africa. In order to comply with the HEQSF, all institutions are legally required to participate in a national initiative led by regulatory bodies such as the Department of Higher Education and Training (DHET), the Council on Higher Education (CHE), and the South African Qualifications Authority (SAQA). The University of Pretoria is presently engaged in an ongoing effort to align its qualifications and programmes with the HEQSF criteria. Current and prospective students should take note that changes to UP qualification and programme names, may occur as a result of the HEQSF initiative. Students are advised to contact their faculties if they have any questions.